

8-24 $P_{1s}(r) = \frac{4}{a_0^3} r^2 e^{-2r/a_0}$ for hydrogen ground state, $U(r) = -\frac{ke^2}{r}$ is potential energy ($Z = 1$)

$$\begin{aligned}\langle U \rangle &= \int_0^\infty U(r) P_{1s}(r) dr = -\frac{4ke^2}{a_0^3} \int_0^\infty r e^{-2r/a_0} dr \\ &= -\frac{4ke^2}{a_0^3} \left(\frac{a_0}{2}\right)^2 \int_0^\infty z e^{-z} dz \quad \text{where } z = \frac{2r}{a_0} \\ &= \frac{-ke^2}{a_0} = -2(13.6 \text{ eV}) = -27.2 \text{ eV}.\end{aligned}$$

To find $\langle K \rangle$, we note that $\langle K \rangle + \langle U \rangle = \langle E \rangle = -\frac{ke^2}{2a_0} = -13.6 \text{ eV}$ so, $\langle K \rangle = \frac{ke^2}{2a_0} = +13.6 \text{ eV}$.

8-25 The most probable distance is the value of r which maximizes the radial probability density $P(r) = |rR(r)|^2$. Since $P(r)$ is largest where $rR(r)$ reaches its maximum, we look for the most probable distance by setting $\frac{d\{rR(r)\}}{dr}$ equal to zero, using the functions $R(r)$ from Table 8.4. For clarity, we measure distances in bohrs, so that $\frac{r}{a_0}$ becomes simply r , etc. Then for the 2s state of hydrogen, the condition for a maximum is

$$0 = \frac{d}{dr} \left\{ (2r - r^2) e^{-r/2} \right\} = \left\{ 2 - 2r - \frac{1}{2} (2r - r^2) \right\} e^{-r/2}$$

or $0 = 4 - 6r + r^2$. There are two solutions, which may be found by completing the square to get $0 = (r - 3)^2 - 5$ or $r = 3 \pm \sqrt{5}$ bohrs. Of these $r = 3 + \sqrt{5} = 5.236 a_0$ gives the largest value of $P(r)$, and so is the most probable distance. For the 2p state of hydrogen, a similar analysis gives $0 = \frac{d}{dr} \left\{ r^2 e^{-r/2} \right\} = \left\{ 2r - \frac{1}{2} r^2 \right\} e^{-r/2}$ with the obvious roots $r = 0$ (a minimum) and $r = 4$ (a maximum). Thus, the most probable distance for the 2p state is $r = 4a_0$, in agreement with the simple Bohr model.

8-26 The probabilities are found by integrating the radial probability density for each state, $P(r)$, from $r = 0$ to $r = 4a_0$. For the 2s state we find from Table 8.4 (with $Z = 1$ for hydrogen)

$$P_{2s}(r) = |rR_{2s}(r)|^2 = (8a_0)^{-1} \left(\frac{r}{a_0}\right)^2 \left(2 - \frac{r}{a_0}\right)^2 e^{-r/a_0} \quad \text{and} \quad P = (8a_0)^{-1} \int_0^{4a_0} \left(\frac{r}{a_0}\right)^2 \left(2 - \frac{r}{a_0}\right)^2 e^{-r/a_0} dr.$$

Changing variables from r to $z = \frac{r}{a_0}$ gives $P = 8^{-1} \int_0^4 (4z^2 - 4z^3 + z^4) e^{-z} dz$. Repeated integration by parts gives

$$\begin{aligned}P &= 8^{-1} \left\{ (4z^2 - 4z^3 + z^4) - (8z - 12z^2 + 4z^3) - (8 - 24z + 12z^2) - (-24 + 24z) - (24) \right\} e^{-z} \Big|_0^4 \\ &= 8^{-1} \left\{ -(64 + 96 + 104 + 72 + 24) e^{-4} + 8 \right\} = 0.176\end{aligned}$$

For the $2p$ state of hydrogen $P_{2p}(r) = |rR_{2p}(r)|^2 = (24a_0)^{-1} \left(\frac{r}{a_0}\right)^4 e^{-r/a_0}$ and

$$P = (24a_0)^{-1} \int_0^{4a_0} \left(\frac{r}{a_0}\right)^4 e^{-r/a_0} dr = 24^{-1} \int_0^4 z^4 e^{-z} dz. \text{ Again integrating by parts, we get}$$

$P = 24^{-1} \{-z^4 - 4z^3 - 12z^2 - 24z - 24\} e^{-z} \Big|_0^4 = 24^{-1} \{-824e^{-4} + 24\} = 0.371$. The probability for the $2s$ electron is much smaller, suggesting that this electron spends more of its time in the outer regions of the atom. This is in accord with classical physics, where the electron in a lower angular momentum state is described by orbits more elliptic in shape.

8-29 To find Δr we first compute $\langle r^2 \rangle$ using the radial probability density for the $1s$ state of

hydrogen: $P_{1s}(r) = \frac{4}{a_0^3} r^2 e^{-2r/a_0}$. Then $\langle r^2 \rangle = \int_0^\infty r^2 P_{1s}(r) dr = \frac{4}{a_0^3} \int_0^\infty r^4 e^{-2r/a_0} dr$. With $z = \frac{2r}{a_0}$, this is

$\langle r^2 \rangle = \frac{4}{a_0^3} \left(\frac{a_0}{2}\right)^5 \int_0^\infty z^4 e^{-z} dz$. The integral on the right is (see Example 8.9) $\int_0^\infty z^4 e^{-z} dz = 4!$ so that

$\langle r^2 \rangle = \frac{4}{a_0^3} \left(\frac{a_0}{2}\right)^5 (4!) = 3a_0^2$ and $\Delta r = \left(\langle r^2 \rangle - \langle r \rangle^2\right)^{1/2} = \left[3a_0^2 - (1.5a_0)^2\right]^{1/2} = 0.866a_0$. Since Δr is an appreciable fraction of the average distance, the whereabouts of the electron are largely unknown in this case.

8-30 The averages $\langle r \rangle$ and $\langle r^2 \rangle$ are found by weighting the probability density for this state

$P_{1s}(r) = 4\left(\frac{Z}{a_0}\right)^3 r^2 e^{-2Zr/a_0}$ with r and r^2 , respectively, in the integral from $r = 0$ to $r = \infty$:

$$\langle r \rangle = \int_0^\infty r P_{1s}(r) dr = 4 \left(\frac{Z}{a_0}\right)^3 \int_0^\infty r^3 e^{-2Zr/a_0} dr$$

$$\langle r^2 \rangle = \int_0^\infty r^2 P_{1s}(r) dr = 4 \left(\frac{Z}{a_0}\right)^3 \int_0^\infty r^4 e^{-2Zr/a_0} dr$$

Substituting $z = \frac{2Zr}{a_0}$ gives

$$\langle r \rangle = 4 \left(\frac{Z}{a_0}\right)^3 \left(\frac{a_0}{2Z}\right)^4 \int_0^\infty z^3 e^{-z} dz = \frac{3!}{4} \left(\frac{a_0}{Z}\right) = \frac{3}{2} \left(\frac{a_0}{Z}\right)$$

$$\langle r^2 \rangle = 4 \left(\frac{Z}{a_0}\right)^3 \left(\frac{a_0}{2Z}\right)^5 \int_0^\infty z^4 e^{-z} dz = \frac{4!}{8} \left(\frac{a_0}{Z}\right)^2 = 3 \left(\frac{a_0}{Z}\right)^2$$

and $\Delta r = \left(\langle r^2 \rangle - \langle r \rangle^2\right)^{1/2} = \frac{a_0}{Z} \left[3 - \frac{9}{4}\right]^{1/2} = 0.866 \left(\frac{a_0}{Z}\right)$. The momentum uncertainty is deduced from the average potential energy

$$\langle U \rangle = -kZe^2 \int_0^\infty \frac{1}{r} P_{1s}(r) dr = -4kZe^2 \left(\frac{Z}{a_0}\right)^3 \int_0^\infty r e^{-2Zr/a_0} dr = -4kZe^2 \left(\frac{Z}{a_0}\right)^3 \left(\frac{a_0}{2Z}\right)^2 = -\frac{k(Ze)^2}{a_0}.$$

Then, since $E = -\frac{k(Ze)^2}{2a_0}$ for the 1s level, and $a_0 = \frac{\hbar^2}{m_e k e^2}$, we obtain

$$\langle p^2 \rangle = 2m_e \langle K \rangle = 2m_e (E - \langle U \rangle) = \frac{2m_e k (Ze)^2}{2a_0} = \left(\frac{Z\hbar}{a_0} \right)^2.$$

With $\langle \mathbf{p} \rangle = 0$ from symmetry, we get $\Delta p = \left(\langle p^2 \rangle \right)^{1/2} = \frac{Z\hbar}{a_0}$ and $\Delta r \Delta p = 0.866\hbar$ for any Z , consistent with the uncertainty principle.

9-1

$$\Delta E = 2\mu_B B = hf$$

$$2(9.27 \times 10^{-24} \text{ J/T})(0.35 \text{ T}) = (6.63 \times 10^{-34} \text{ Js})f \text{ so } f = 9.79 \times 10^9 \text{ Hz}$$

9-4

(a) $3d$ subshell $\Rightarrow l=2 \Rightarrow m_l = -2, -1, 0, 1, 2$ and $m_s = \pm \frac{1}{2}$ for each m_l

	l	m_l	m_s
n			
3	2	-2	-1/2
3	2	-2	+1/2
3	2	-1	-1/2
3	2	-1	+1/2
3	2	0	-1/2
3	2	0	+1/2
3	2	1	-1/2
3	2	1	+1/2
3	2	2	-1/2
3	2	2	+1/2

(b) $3p$ subshell: for a p state, $l=1$. Thus m_l can take on values $-l$ to l , or $-1, 0, 1$. For each m_l , m_s can be $\pm \frac{1}{2}$.

	l	m_l	m_s
n			
3	1	-1	-1/2
3	1	-1	+1/2
3	1	0	-1/2
3	1	0	+1/2
3	1	1	-1/2
3	1	1	+1/2

9-6

The exiting beams differ in the spin orientation of the outermost atomic electron. The energy difference derives from the magnetic energy of this spin in the applied field \mathbf{B} :

$$U = -\boldsymbol{\mu}_s \cdot \mathbf{B} = g \left(\frac{-e}{2m} \right) S_z B = -g\mu_B B m_s.$$

With $g = 2$ for electrons, the energy difference between the up spin ($m_s = \frac{1}{2}$) and down spin ($m_s = -\frac{1}{2}$) orientations is

$$\Delta U = g\mu_B B = (2)(9.273 \times 10^{-24} \text{ J/T})(0.5 \text{ T}) = 9.273 \times 10^{-24} \text{ J} = 5.80 \times 10^{-5} \text{ eV}.$$

9-17 From Equation 8.9 we have $E = \left(\frac{\hbar^2 \pi^2}{2mL^2}\right)(n_1^2 + n_2^2 + n_3^2)$

$$E = \frac{(1.054 \times 10^{-34})^2 (\pi^2)(n_1^2 + n_2^2 + n_3^2)}{2(9.11 \times 10^{-31})(2 \times 10^{-10})^2} = (1.5 \times 10^{-18} \text{ J})(n_1^2 + n_2^2 + n_3^2) = (9.4 \text{ eV})(n_1^2 + n_2^2 + n_3^2)$$

(a) 2 electrons per state. The lowest states have

$$(n_1^2 + n_2^2 + n_3^2) = (1, 1, 1) \Rightarrow E_{111} = (9.4 \text{ eV})(1^2 + 1^2 + 1^2) \text{ eV} = 28.2 \text{ eV}.$$

For $(n_1^2 + n_2^2 + n_3^2) = (1, 1, 2)$ or $(1, 2, 1)$ or $(2, 1, 1)$,

$$E_{112} = E_{121} = E_{211} = (9.4 \text{ eV})(1^2 + 1^2 + 2^2) = 56.4 \text{ eV}$$

$$E_{\min} = 2 \times (E_{111} + E_{112} + E_{121} + E_{211}) = 2(28.2 + 3 \times 56.4) = 398.4 \text{ eV}$$

(b) All 8 particles go into the $(n_1^2 + n_2^2 + n_3^2) = (1, 1, 1)$ state, so

$$E_{\min} = 8 \times E_{111} = 225.6 \text{ eV}.$$

9-21 (a) $1s^2 2s^2 2p^4$

(b) For the two 1s electrons, $n = 1, l = 0, m_l = 0, m_s = \pm \frac{1}{2}$.

For the two 2s electrons, $n = 2, l = 0, m_l = 0, m_s = \pm \frac{1}{2}$.

For the four 2p electrons, $n = 2, l = 1, m_l = 1, 0, -1, m_s = \pm \frac{1}{2}$.

9-15 The spin of the atomic electron has a magnetic energy in the field of the orbital moment given by Equations 9.6 and 9.12 with a g -factor of 2, or $U = -\boldsymbol{\mu}_s \cdot \mathbf{B} = 2\left(\frac{e}{2m_e}\right)S_z B = 2\mu_B m_s B$.

The magnetic field \mathbf{B} originates with the orbiting electron. To estimate \mathbf{B} , we adopt the equivalent viewpoint of the atomic nucleus (proton) circling the electron, and borrow a result from classical electromagnetism for the \mathbf{B} field at the center of a circular current

loop with radius r : $B = \frac{2k_m \mu}{r^3}$. Here k_m is the magnetic constant and $\mu = i\pi r^2$ is the magnetic moment of the loop, assuming it carries a current i . In the atomic case, we identify r with the orbit radius and the current i with the proton charge $+e$ divided by the orbital period $T = \frac{2\pi r}{v}$. Then $\mu = \frac{evr}{2} = \left(\frac{e}{2m_e}\right)L$ where $L = m_e vr$ is the orbital angular

momentum of the *electron*. For a p electron $l=1$ and $L = [l(l+1)]^{1/2}\hbar = \sqrt{2}\hbar$, so

$\mu = \left(\frac{e\hbar}{2m_e}\right)\sqrt{2} = \mu_B\sqrt{2} = 1.31 \times 10^{-23}$ J/T. For r we take a typical atomic dimension, say $4a_0 (= 2.12 \times 10^{-10}$ m) for a $2p$ electron, and find

$$B = \frac{2(10^{-7} \text{ N/A}^2)(1.31 \times 10^{-23} \text{ J/T})}{(2.12 \times 10^{-10} \text{ m})^3} = 0.276 \text{ T}.$$

Since m_s is $\pm \frac{1}{2}$ the magnetic energy of the electron spin in this field is

$$U = \pm \mu_B B = \pm (9.27 \times 10^{-24} \text{ J/T})(0.276 \text{ T}) = \pm 2.56 \times 10^{-24} \text{ J} = \pm 1.59 \times 10^{-5} \text{ eV}.$$

The up spin orientation (+) has the higher energy; the predicted energy difference between the up (+) and down (-) spin orientations is twice this figure, or about 3.18×10^{-5} eV—a result which compares favorably with the measured value, 5×10^{-5} eV.

9-24

Atom	3s	3p			4s	Electron Configuration
Na	\uparrow					[Ne]3s ¹
Mg	$\uparrow\downarrow$					[Ne]3s ²
Al	$\uparrow\downarrow$	\uparrow				[Ne]3s ² 3p ¹
Si	$\uparrow\downarrow$	\uparrow	\uparrow			[Ne]3s ² 3p ²
P	$\uparrow\downarrow$	\uparrow	\uparrow	\uparrow		[Ne]3s ² 3p ³
S	$\uparrow\downarrow$	$\uparrow\downarrow$	\uparrow	\uparrow		[Ne]3s ² 3p ⁴
Cl	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	\uparrow		[Ne]3s ² 3p ⁵
Ar	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$		[Ne]3s ² 3p ⁶
K	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	$\uparrow\downarrow$	\uparrow	[Ar]4s ¹

The 3s subshell is energetically lower and so fills before the 3p. According to Hund's rule, electrons prefer to align their spins so long as the exclusion principle can be satisfied.